

## Camphor Ionic Liquid: Correlation between Stereoselectivity and Cation-Anion Interaction

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CS: camphorsulfonate

As a halogen-free anion for an imidazolium room temperature ionic liquid, the use of a camphorsulfonate causes an increase in the number of free (naked) imidazolium cations, which produces an effective *endolexo* stereoselective Diels—Alder reaction.

Room temperature ionic liquids (ILs) have been widely used as solvents for various reactions due partly to their unique and the environmentally benign properties.  $^1$  Since a Diels—Alder addition can sensitively reflect the microscopic environment of solutes surrounded by medium,  $^2$  the reaction has often been applied for IL.  $^{2b,3}$  The reaction in imidazolium ILs showed that the hydrogen bond between dienophile and the imidazolium cation via the  $C_2$ —H affected the *endolexo* stereoselectivity of the product.  $^{3d}$  In particular, the counterion of the imidazolium cation is known to have an influence on the hydrogen bond.  $^{3d,4}$  That is, an anion acting as a poor hydrogen bond acceptor is supposed to show high ste-

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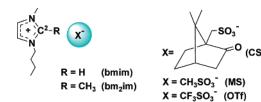
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CHART 1



reoselectivity. However, the question still remains as to what anionic structures of imidazolium IL are most appropriate for effective stereoselectivity.

For the purpose of nonfluorous imidazolium ILs such as sugar type salts,<sup>5</sup> in this research we use a camphor-sulfonate (CS) as one of the halogen-free anions with 1-butyl-3-methylimidazolium (bmim) cation in favor of its bulky shape. This attempt allows us to confirm first, despite its bulkiness and high molecular weight, that the camphor salt ([bmim][CS]) becomes a viscous liquid even on a salted ice bath, while the bmim methanesulfonate ([bmim][MS]) is solid at room temperature.<sup>6</sup> In this study, we describe the preparation and physicochemical properties of the camphor ILs with a focus on the endo/exo stereoselectivity of a Diels—Alder reaction in ILs. The reactivity and the liquescency are to be viewed from the microscopic environment of ions, i.e., ion states such as contact ion pair and free ions.

Adopting the anion-exchange resin method for [bmim]-[Br] and [bm2im][Cl] with chiral camphor-10-sulfonic acid produced a pale yellow liquid [bmim][CS] and a solid [bm<sub>2</sub>im][CS], respectively, with quantitative yields regardless of its chirality. The ion exchanged ratios for both salts were ascertained to be above 98.5% by <sup>1</sup>H NMR spectra. The aqueous solutions of (S)-(+)-[bmim][CS] and (R)-(-)-[bmim][CS] showed symmetrical circular dichroism spectra and specific rotations. The [bmim] [CS] was miscible with acetone and protic solvents such as water and alcohols, and immiscible with hexane, ether, and ethyl acetate. The [bmim][CS]  $(1.9 \times 10^4 \text{ cP at } 25 \text{ °C})$ showed 229 times the viscosity of the ordinary IL, [bmim]- $[BF_4]$  (83 cP at 25 °C). The liquid of (S)-(+)-[bmim] [CS] containing water by 0.04 wt % and the solid of (S)-(+)-[bm<sub>2</sub>im][CS] were used for the following experiments.

Since the undiluted [bmim] [CS] is too viscous to use as a reaction solvent, we mixed it with a fluid IL for a Diels—Alder reaction. The reaction between cyclopentadien and ethylacrylate was conducted without any catalyst in the [Im] [CS] (Im = bmim, bm<sub>2</sub>im) diluted with [bmim] [BF<sub>4</sub>] by 15/100 molar ratio. As the control experiments, the reactions were also carried out in the mixture of [bmim] [X] (X = MS, OTf) and [bmim] [BF<sub>4</sub>] by the same molar ratio as the [bmim] [CS]. Even after addition of diene and dienophile, all the reaction mixtures retained a single homogeneous phase. The endo/exo ratios of the isolated products were calculated by the <sup>1</sup>H NMR spectra.

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<sup>(6)</sup> Nobuoka, K. Abstracts chemwest 2003 (Japan), 2003, No. 84. (7) (S)-(+)-[bmim][CS]:  $+1700 \Delta E$  at 290 nm,  $[\alpha]^{25}_{589} +7.9$ ; (R)-(-)-[bmim][CS]:  $-1700 \Delta E$  at 290 nm,  $[\alpha]^{25}_{589} -7.6$ .

TABLE 1. Effect of Anion of IL on the Stereoselectivity of Diels-Alder Reaction in ILs

added sulfonate in [bmim][BF <sub>4</sub> ]	<i>T</i> /°C	$[endo]/[exo]^a$	isolated yield/%
none [bmim][MS] $^b$ [bmim][OTf] $^b$ [bmim][CS] $^b$	20	3.4 4.1 4.0 6.1	36 42 16 29
none [bmim] [CS] $^b$	-10	6.1 10.3	60 66
$[\mathrm{bm}_2\mathrm{im}][\mathrm{CS}]^b$	20	3.0	28

 $^a$  Calculated from  $^1$ H NMR spectra, [diene] = 2.0 M, [dienophile] = 1.3 M, reaction time 20 h.  $^b$  The molar ratio of [bmim][X]/[bmim][BF<sub>4</sub>] is 15/100 (X = MS, OTf, CS).

Table 1 shows the endo/exo ratio of the isolated product in the ILs. For the bmim ILs, the pure [bmim]-[BF<sub>4</sub>] at 20 °C allowed the reaction to show the lowest endo/exo value to be 3.4, which is comparable to that of a diethyl ether.<sup>8</sup> On the other hand, adding the sulfonates heightened the value by 0.7, 0.6, and 2.7 for [bmim][MS], [bmim][OTf], and [bmim][CS], respectively. The added [bmim][CS] nearly doubled the stereoselectivity of the pure [bmim][BF<sub>4</sub>] solvent. The *endo* form is more favorable in the IL with bulky [CS] than with small [MS] or [OTf]. Reducing the reaction temperature down to -10°C for the [bmim][BF<sub>4</sub>] with and without [bmim][CS] improved both the isolate yields and the endo selectivity to be around 60% and above 6.1, respectively. The addition of [bmim] [CS] at -10 °C resulted in an increase in the endolexo stereoselectivity to 10.3 similar to the reaction done at 20 °C. These findings provide evidence that the endo selectivity is attributed to the presence of the [bmim][CS] regardless of the reaction temperature. Such high stereoselectivity shows the endolexo value to be 10 or over, which was reported only in the cases where catalysts were used.3g As far as we know, however, no such excellent stereoselectivity to achieve the value to be as much as 10.3 in ILs, without using catalyst, has been reported heretofore.

As aforementioned, the use of the CS anion facilitates the endo selective reaction when the cationic module is bmim. On the contrary, when  $[bm_2im][CS]$  is applied in place of [bmim][CS], the endo/exo ratio is as little as 3.0. This value is 0.4 less than that of the pure  $[bmim][BF_4]$ . Unlike bmim, the  $bm_2im$  ring lacks the  $C_2-H$  by which the hydrogen bond between dienophile and the imidazolium ring can be formed. These findings, therefore, indicate that the presence of  $C_2-H$  accompanied by a CS anion is instrumental for the effective  $\mbox{\it endo/exo}$  selectivity.

Changing the anion of a salt affects the microenvironment of ions. In this study we evaluated the cation—anion association in the ILs by means of the specially designed ES mass spectrometer with positively charged mode. In this mass spectrometry, fragmentation of emerged liquid

droplets via adiabatic expansion in a vacuum chamber generates molecular clusters of the ILs. Unless the interaction between cation(s) and anions(s) is strong enough to be detected as clusters, the emerged clusters eventually become vaporized as monomeric ions. Therefore, a comparison of the cluster peaks of the ion pair with the monomeric peaks allows us to estimate the cation—anion interaction.

Figure S5 (see the Supporting Information) shows the mass spectra measured for the methanol solutions of the sulfonate ILs. To obtain efficiently liquid droplets by the electrospray, the ILs were dissolved in methanol. The detected clusters from the CS and other sulfonate salts were found to be very different. The [bmim][MS] and [bmim][OTf] both provide the nonsolvated imidazolium free cation, [bmim]+, and the nonsolvated clusters composed of the ion pairs ranging from monomer (n = 0) to heptamer (n = 7),  $[bmim]^+([bmim]^+[X]^-, X = MS, OTf)_n$ . On the contrary, the [bmim][CS] shows predominantly the nonsolvated imidazolium free cations which never produce such clusters of ion pairs except in the case of the monomeric ion pairs, [bmim]<sup>+</sup>([bmim]<sup>+</sup>[CS]<sup>-</sup>)<sub>1</sub>. Even the monomeric ion pair of CS salt was outnumbered by that of MS or OTf salts. The same phenomenon as the [bmim][CS] can be observed for the methanolic [bm2im]-[CS]. The difference in the clustering indicates that the bulky [CS] anion binds more loosely with the imidazolium cation than the small [MS] or [OTf] anions. As a result, [bmim][CS] and [bm2im][CS] both provide free (naked) imidazolium cation more efficiently than [bmim][MS] or [bmim][OTf].

Unlike inorganic salts such as LiCl,<sup>9</sup> the organic ILs show almost no solvated peaks. 10 Hence, the observed difference in the ion states of the methanolic ILs could be adopted for the corresponding undiluted ILs. In fact, <sup>1</sup>H NMR spectra of the undiluted ILs revealed a weak interaction between imidazolium and the bulky [CS] as well as the ES-MS analysis of the methanolic solutions. The undiluted [bmim][CS] exhibits the lowest chemical shift of the C2 proton at 8.68 ppm among a series of undiluted [bmim][X], namely, the chemical shifts in ppm at 30 °C for [BF<sub>4</sub><sup>-</sup>], [TFSI<sup>-</sup>], <sup>3d</sup> [CF<sub>3</sub>SO<sub>3</sub><sup>-</sup>], <sup>3d</sup> and [CF<sub>3</sub>CO<sub>2</sub><sup>-</sup>] <sup>3d</sup> are 8.84, 8.70, 9.00, and 9.80, respectively. The lower value than [bmim][BF4], whose anion acts as a poor hydrogen bond acceptor, means that the bulky [CS] anion is better suited for forcing the cation to be free from hydrogen bonds than the other anions.

The calculated molecular volume for CS (196 ų) is 3.4 and 2.9 times larger than that for MS (57 ų) and OTf (68 ų), respectively. Among three sulfonates, the largest CS anion could place the cationic bmim as far away as possible from itself, increasing naked bmims. It is reasonable for us to conclude that the large CS anion impairs the molecular packing of the salt to the point of becoming liquid and causes an increase in the number

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of free cations. This facilitates the formation of the hydrogen bond between the  $C_2-H$  of the cation and the carbonyl of the dienophile. Like a Lewis acid catalyst for the dienophile,  $^{12}$  the hydrogen bond can also stabilize the LUMO energy of the dienophile, providing the *endo* isomer as the preferred species.  $^{3d,13}$ 

In conclusion, we have attempted to demonstrate that utilizing the CS anion makes the imidazolium cation free enough to expedite the stereoselectivity compared to MS and OTf anions. It can be deduced that the reactivity is relevant to the cation—anion association, i.e., ion states of the imidazolium ILs at least for CS, MS, and OTf anions. To determine whether devising the size of the anions is the effective methodology for regulating the ion states of ILs, various sulfonates, which are different in size, are currently coupled with bmim.

## **Experimental Section**

Preparation of 3-Butyl-1-methylimidazolium Camphorsulfonate ([bmim][CS]). [Bmim][Br] prepared by a general  $\mathrm{method}^{14}\,\mathrm{was}\,\mathrm{transformed}$  into the aimed camphor sulfonate salt via [bmim]<sub>2</sub>[SO<sub>4</sub>] by means of a strong base anion exchange column chromatography. Into a 400 mL aqueous solution of [bmim][Br] (39 g, 0.18 mol) was added concentrated H<sub>2</sub>SO<sub>4</sub> (5 mL) with stirring for 3 h at room temperature. After removal of the solvent, the residual liquid was further dried in vacuo (under 0.1 mbar) at 60 °C for 24 h to remove water and HBr as much as possible. [Bmim]<sub>2</sub>[SO<sub>4</sub>] (37 g) was obtained as a pale yellow viscous oil. The obtained oil was subjected to the column packed with the resin, which was replaced by camphorsulfonate to give pale yellow viscous oil. The oil was dried in vacuo (under 0.1 mbar) at 60 °C for 5 days, producing a pale yellow viscous liquid (yield: 64.8 g, 97%). The ion exchanged ratio was calculated by the <sup>1</sup>H NMR spectrum in Figure S1, i.e., the ratio of the average of the integrated values of two  $CH_3$ s of [CS] at 1.03 and 0.73 ppm (3.10) against the integrated values of CH<sub>2</sub>CH<sub>3</sub> of [bmim] at 0.89 ppm (3.15) was used.  $^{1}$ H NMR (300 MHz, DMSO- $d_6$ -TMS):  $\delta$  9.12 (s, 1H, C(2)H), 7.76 (t, J = 1.7 Hz, 1H, C(4)H), 7.69 (t, J = 1.7 Hz, 1H, C(5)H), 4.15 (t, J = 7.2 Hz, 2H, N-C $H_2$ ), 3.84 (s, 3H,  $N-CH_3$ ), 2.89 (d, J = 14.7 Hz, 1H, CS), 2.65 (m, 1H, CS), 2.39 (d, J = 14.7 Hz, 1H, CS), 2.22 (tt, J = 18.1 Hz, J= 3.8 Hz, 1H, CS), 1.94 (t, J = 4.4 Hz, 1H, CS), 1.93 - 1.72 (m,4H, N-CH<sub>2</sub>CH<sub>2</sub> and CS), 1.30-1.21 (m, 4H, CH<sub>2</sub>CH<sub>3</sub> and CS),  $1.03 \text{ (s, 3H, CS}-\text{C}H_3), 0.89 \text{ (t, } J = 7.3 \text{ Hz, 3H, CH}_2\text{C}H_3), 0.73 \text{ (s, }$ 3H, CS-C $H_3$ ). <sup>13</sup>C NMR (75.46 MHz, DMSO- $d_6$ -TMS):  $\delta$  216.4 (s), 137.1 (t), 124.4–122.0 (m, 2C), 58.4 (s), 48.7–46.8 (m, 3C), 43.4-41.8 (m, 2C), 35.8 (m), 32.9-30.1 (m), 27.3-23.4 (m, 2C), 20.6–18.1 (m, 3C), 13.4 (q, J = 65.7). UV-vis (H<sub>2</sub>O)  $\lambda_{max}$ /nm  $(\log \epsilon)\, 265\, (4.04).\, Anal.\, Calcd\, for\, C_{18}H_{30}N_2O_4S:\,\, C,\, 58.35;\, H,\, 8.16;$ N, 7.56. Found: C, 58.17; H, 8.43; N, 7.74.

**Preparation of 3-Butyl-1,2-dimethylimidazolium Camphorsulfonate** ([bm2im][CS]). [Bm2im][CS] was prepared by using the same procedure as described above in the preparation of [bm2im][CS] from [bm2im][Br] and camphorsulfonic acid (yield: 4.3 g, 85%). [Bmim][CS] is a viscous liquid, whereas [bm2im][CS] is a white solid. <sup>1</sup>H NMR (300 MHz, DMSO- $d_6$ –TMS):  $\delta$  7.73 (d, J = 2.1 Hz, 1H, C(4)H), 7.70 (d, J = 2.1 Hz, 1H, C(5)H), 4.17 (t, J = 7.3 Hz, 2H, N- $CH_2$ ), 3.81 (s, 3H, N- $CH_3$ ), 2.91 (d, J = 14.7 Hz, 1H, CS), 2.75 (m, H, CS), 2.65 (S,3H, C- $CH_3$ ), 2.41 (d, J = 14.7 Hz, 1H, CS), 2.28 (tt, J = 18.1 Hz, J = 3.6 Hz, 1H, CS), 1.99 (t, J = 4.5 Hz, 1H, CS), 1.96–1.72 (m, 4H, N- $CH_2CH_2$  and CS), 1.40–1.25 (m, 4H, C $H_2CH_3$  and

CS), 1.10 (s, 3H, CS- $CH_3$ ), 0.95 (t, J=7.3 Hz, 3H,  $CH_2CH_3$ ), 0.79(s, 3H,  $CS-CH_3$ ).  $^{13}C$  NMR (75.46 MHz, DMSO- $d_6$ -TMS):  $\delta$  216.4 (s), 144.4 (s), 123.1–120.5 (m, 2C), 58.4 (s), 47.4–46.4 (m, 3C), 43.3–41.8 (m, 2C), 34.8 (m), 31.4 (m), 27.3–23.3 (m, 2C), 21.6–18.3 (m, 3C), 13.5 (q, J=65.7 Hz), 9.3 (m, J=105.3 Hz). Anal. Calcd for  $C_{19}H_{32}N_2O_4S$ : C, 59.35; H, 8.39; N, 7.29. Found: C, 59.07; H, 8.47; N, 7.29.

Diels-Alder Reaction in ILs. Cyclopentadiene (0.37 mL, 4.5 mmol) was dissolved in [bmim][BF<sub>4</sub>] (2.3 mL) or in a mixture of [bmim][CS], [bmim][MS], [bmim][OTf] or [bm2im][CS], and [bmim][BF<sub>4</sub>] (15:100 mol/mol, 2.3 mL). Ethyl acrylate (0.30 mL, 3.0 mmol) was added into the diene solution, which was then stirred for 20 h at room temperature or −10 °C. The reaction mixture maintained a single phase, not biphase. Extraction with diethyl ether followed by purification column chromatography (silicagel, BW-80S, Fuji Silysia, ethyl acetate/ hexane 1/10 (v/ v)) provided the isolate Diels-Alder products. The endolexo ratios were calculated by means of <sup>1</sup>H NMR spectra of Diels-Alder products. The CHCO<sub>2</sub>C<sub>2</sub>H<sub>5</sub> proton in DMSO-d<sub>6</sub> appears at 3.20 and 3.03 ppm as endo and exo isomers, respectively. These signals were used for the determination of the endolexo ratios. All ionic liquids were vacuum dried (0.08 mmHg, 60 °C, 5days) prior to use.

**Positive Ion Mass spectra.** To obtain information on the microscopic structures of ions, we have used the mass spectroscopy designed specially for observing clusters isolated through fragmentation of liquid droplets. For the analysis of electrolyte solution, an electrospray interface was used to form the liquid droplets. The details of this procedure have been reported previously.8 As is shown schematically in the Supporting Information, the mass spectrometer is composed of a five-stage differentially pumped vacuum system, a homemade electrospray interface and a quadrupole mass filter (Extral C50). For the electrospray, electric voltages were supplied to the nozzle and three skimmers  $(E_1-E_4)$ , and nitrogen gas was made to flow to maintain an appropriate pressure balance  $(P_1-P_5)$ . Positive ion mass spectra of clusters generated from methanolic ILs at room temperature include the following: (a) [bmim][MS] (electric voltage:  $E_1 = +2.84 \text{ kV}, E_2 +208 \text{ V}, E_3 = +211 \text{ V}, E_4 = +178 \text{ V},$ pressure:  $P_1 = 647.8 \text{ Torr}, P_2 = 9.58 \text{ Torr}, P_3 = 0.007 \text{ Torr}, P_4 = 0.007 \text{ Torr}$ = 1.13  $\times$  10<sup>-5</sup> Torr,  $P_5$  = 6.30  $\times$  10<sup>-7</sup> Torr); (b) [bmim][OTf] (electric voltage:  $E_1 = +2.87 \text{ kV}, E_2 = +207 \text{ V}, E_3 = +211 \text{ V}, E_4$ = +177 V; pressure:  $P_1$  = 646.1 Torr,  $P_2$  = 9.63 Torr,  $P_3$  = 0.007 Torr,  $P_4 = 1.14 \times 10^{-5}$  Torr,  $P_5 = 8.90 \times 10^{-7}$  Torr); (c) [bmim]-[CS] (electric voltage:  $E_1 = +3.06$  kV,  $E_2 = +206$  V,  $E_3 = +218$  $V, E_4 = +178 V;$  pressure:  $P_1 = 646.1 \text{ Torr}, P_2 = 9.81 \text{ Torr}, P_3 = 9.81 \text{ Torr}$ 0.008 Torr,  $P_4 = 1.23 \times 10^{-5}$  Torr,  $P_5 = 8.30 \times 10^{-7}$  Torr); and (d) [bm<sub>2</sub>im][CS] (electric voltage:  $E_1 = +4.61 \text{ kV}, E_2 = +207 \text{ V},$  $E_3 = +218 \text{ V}, E_4 = +177 \text{ V}; \text{ pressure: } P_1 = 645.7 \text{ Torr}, P_2 = 9.83$ Torr,  $P_3 = 0.008$  Torr,  $P_4 = 1.01 \times 10^{-5}$  Torr,  $P_5 = 3.34 \times 10^{-7}$ Torr).

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**Supporting Information Available:** General methods, <sup>1</sup>H NMR spectra of [bmim][CS] and [bm<sub>2</sub>im][CS], UV—vis and CD spectra of [bmim][CS], a schematic illustration of the mass spectrometry for clusters isolated from electrolyte solutions with ES-MS, and a set of positive ion mass spectra of clusters generated from methanolic ILs. This material is available free of charge via the Internet at http://pubs.acs.org.

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